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number of protons in the nucleus increase; however the distance between the nucleus and the valence electrons remains the same. Thus the protons have a stronger pull on the valence electrons decreasing the size of the atom.

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molar mass of $C_6H_6O = 94.11 \text{ g}$ 21. a. molar mass of $SO_3 = 80.07 \text{ g}$ 49.2 mg SO_3 1 g $\frac{1000 \text{ mg}}{1 \text{ g}}$ 1 mol $80.07 \text{ g} = 6.14 \cdot 10^{-4}$

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mol SO₃ b. molar mass of PbO₂ = 239.2 g 7.44 x 10⁴ kg PbO 2
1000 g 1 kg 1 mol 239.2 g = 3.11 10⁵ mol PbO 2 c. molar mass
of CHCl₃ = 119.37 g 59.1 g CHCl₃ 1 mol 119.37 g = 0.495 mol
CHCl₃ d. molar mass of C₂H₃Cl₃ = 133.39 g 3.27 ...

Chapter 6 Chemical Composition - Francis Howell High School

Chemistry Chapter 6. periodic law. metals. nonmetals. metalloid.
when the elements are arranged in order of increasing atomic
n.... one of a class of elements that are good conductors of heat
an.... an element that tends to be a poor conductor of heat and
elect....

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CHAPTER 6 SOLUTIONS MANUAL 4. Determine why producing 1 million fran-cium atoms per second is not enough to make measurements, such as density of boiling point. Even one million atoms collected together as a solid are microscopic. A grain of salt contains about 10^{15} sodium atoms. Section 6.1 Assessment page 181 1. Describe the development of the modern

The Periodic Table and Periodic LawThe Periodic Table and ...

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lanthanides 10. 2 11. fourth 12. transition elements 13. 32 14.
valence electrons 15. electron affinity 16. electronegativity 17.
ionization energy 18. 3 s² 3p⁴ 19 ...

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Chemistry Chapter 6. Depletion. inherent. Mole. exponents. less in quantity, amount going down. A natural part of something. unit of amount in chemistry. The number that is small and raised to show how many times to....

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1. octaves
2. eight
3. nine
4. accepted
5.

Chapter 6 Test B The Periodic Table Answers

Atoms Chemistry is the study of matter Chemistry in biology
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Chapter 6 Chemistry Test Answer Key chapter 6 chemistry test answer CHAPTER 6 THERMOCHEMISTRY 162 CHAPTER 6: THERMOCHEMISTRY To convert the answer to joules, we write: $1013 \text{ J} \cdot 018 \text{ L atm} = 1 \text{ L atm} = - \cdot \times = \cdot w - 18 \text{ J}$ 617 An expansion implies an increase in volume, therefore w must be -325 J (see the defining equation for pressure-volume

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Answer to 12. Each period in the periodic table corresponds to .

a. an energy level c. an orbital b. an electron d. a suborbital

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